

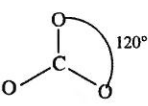
VSEPR Theory:

Shapes:

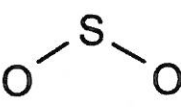
2 e⁻ Domains

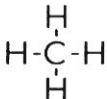
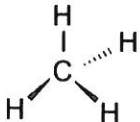
1) Lewis Structure CO ₂	$\ddot{\text{O}}=\text{C}=\ddot{\text{O}}$
2) Shape Name	Linear
3) Shape Drawing	$\text{O}-\text{C}-\text{O}$
4) Bond Angle	180°
5) Hybridization	sp

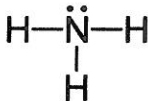
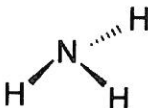
3 e⁻ Domains

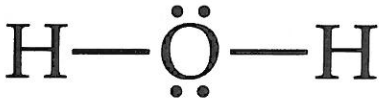
1) Lewis Structure CO ₃ ²⁻	$\left[\begin{array}{c} \ddot{\text{O}}: \\ : \ddot{\text{O}}-\text{C}=\ddot{\text{O}}: \end{array} \right]^{-}$
2) Shape Name	Trigonal Planar
3) Shape Drawing	
4) Bond Angle	120°
5) Hybridization	sp ²

3 e⁻ Domains

1) Lewis Structure SO ₂	$\ddot{\text{O}}=\ddot{\text{S}}-\ddot{\text{O}}:$ (should be in a straight line here)
2) Shape Name	Bent
3) Shape Drawing	
4) Bond Angle	120°
5) Hybridization	sp ²

4 e ⁻ Domains	
1) Lewis Structure CH ₄	
2) Shape Name	Tetrahedral
3) Shape Drawing	
4) Bond Angle	109.5°
5) Hybridization	sp ³

4 e ⁻ Domains	
1) Lewis Structure NH ₃	
2) Shape Name	Trigonal Pyramidal
3) Shape Drawing	
4) Bond Angle	107.5°
5) Hybridization	sp ³

4 e ⁻ Domains	
1) Lewis Structure H ₂ O	
2) Shape Name	Bent
3) Shape Drawing	
4) Bond Angle	105.5°
5) Hybridization	sp ³